## OXIDATION OF 3d TRANSITION METALS BY MoF<sub>6</sub> AND WF<sub>6</sub> IN ACETONITRILE AND SOME REACTIONS OF THE SOLVATED CATION PRODUCTS

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One way of generating solvated metal cations in acetonitrile is to oxidize the metal with MoF<sub>6</sub> or WF<sub>6</sub>. In many cases these reactions are rapid and are thus good synthetic routes to the cations  $[M(NCMe)_{6}]^{2+}$ , M = Fe, Co, Paradoxically, Cu and Zn. The oxidation of nickel is an exception however. massive nickel (powder or wire 99.99%) reacts with WF6 in MeCN at room temperature but not with MoF6, although the latter is the stronger oxidizing MoF, does react when degassed nickel is evaporated in vacuo onto a agent. Pyrex surface but under these conditions no reaction is observed using  $WF_{6}$ . A possible explanation for these observations is that the oxide film on Ni passivates the metal to MoF<sub>6</sub> reaction, while it can be partially removed using the stronger Lewis acid WF<sub>6</sub>. The reduction products in these reactions are  $MoF_6$  or mixtures of  $WF_6$  and  $WF_7$ ; the oxidation product is  $[Ni(NCMe)_6]^{2+}$ . Generation of  $[M(NCMe)_6]^{2+}$  cations under strictly anhydrous conditions allows reactions to be performed which would be difficult or impossible with water [Fe(NCMe)<sub>6</sub>]<sup>2+</sup> forms high-spin [Fe(NCMe)<sub>5</sub>(NMe<sub>3</sub>)]<sup>2+</sup> with NMe<sub>3</sub> in MeCN present. This undergoes stepwise substitution reactions with P(OMe), giving, finally, low-spin  $[Fe(NMe_2){P(OMe)_2}]^{2+}$ . Steric factors appear to be important, as the intermediate steps differ from those observed previously between [Fe(NCMe)<sub>6</sub>]<sup>2+</sup> and P(OMe)<sub>3</sub>. In contrast, NMe<sub>3</sub> is oxidized by [Cu(NCMe)<sub>6</sub>]<sup>2+</sup>, probably via an intermediate  $[Cu(NCMe)_{6-x}(NMe_3)_x]^{2+}$ , and similar behaviour is observed between  $[Cu(NCMe)_{\beta}]^{2+}$  and  $Me_{2}S$ .

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